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Extraction of Ga and Ge from zinc refinery residues in H$_2$C$_2$O$_4$
solutions containing H$_2$O$_2$

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Abstract: A new oxalic acid (H$_2$C$_2$O$_4$) and hydrogen peroxide (H$_2$O$_2$) based extraction process was developed to recover valuable metals - in particular Ga and Ge - from the zinc refinery residues. H$_2$C$_2$O$_4$ and H$_2$O$_2$ leaching studies indicated that the selective leaching of metals in the residues, that primarily contain Zn, Cu, Fe, SiO$_2$, Ga and Ge, can be achieved. Under the optimal leaching conditions ([H$_2$C$_2$O$_4$] = 110 g/L, [H$_2$O$_2$] = 0.12 mol/L, L/S ratio = 8, T = 40 °C and t = 30 min), 99.32% of the Ga, 98.86% of the Ge and 30.25% of the Fe were leached out, whereas the leaching of Zn, Cu and Si only reached 0.30%, 0.82% and 0.43%, respectively. Concerning leachate purification, 98.31% of the iron could be removed in the form of FeC$_2$O$_4$·2H$_2$O with minor losses of Ga and Ge (1.08% and 0.68%) using an ultrasound-assisted iron powder replacement method under optimal conditions. The Ga (99.36%) and Ge (99.89%) were subsequently extracted by tri(octyl-decyl)amine (N235). After extraction, the loaded organic phase was stripped of Ga (98.91%) and Ge (99.21%) into separate solutions using sequential treatments of 2 mol/L H$_2$SO$_4$ and 4 mol/L NaOH respectively. Based on these results, a process flow sheet of efficient separation and recovery of Ga and Ge is presented.

Keywords: Zinc refinery residue; Gallium; Germanium; Oxalic acid; Hydrogen peroxide; Separation
1 Introduction

Owing to their use in technological and strategic applications like semiconductors, infrared optics, optical fibers, medicine and catalysis, gallium and germanium are becoming increasingly important for global industry (Schimmel et al., 2001; Tyszczuk et al., 2007; Depuydt et al., 2006; Fujiwara et al., 2003). Although gallium and germanium are often found together in nature, they generally only exist in trace amounts in many minerals, from which the extraction is not usually economically viable. Currently, gallium and germanium are mainly recovered from the metallurgical by-products of zinc, copper and aluminum refining.

The zinc sulfide concentrate that is produced from a lead and zinc mine in southern China is rich in gallium and germanium; their respective contents are both greater than 100 g/t. The concentrate is leached in a two-stage countercurrent oxygen pressure leaching process. The first stage is performed under high acidity and low temperature and oxygen pressure conditions to leach more gallium and germanium. After the leach solution is pre-neutralized with zinc calcine, zinc powder is added to the solution to replace the gallium and germanium. The resulting material - known as zinc refinery residue - usually contain between 0.2 - 0.5 wt.% Ge and 0.3 - 0.4 wt.% Ga with Zn, SiO$_2$, Cu, Fe and Pb as the main components. As Ga and Ge are difficult to recover and separate from such compounds, currently there is no industrial process for Ge and Ga recovery from these types of residue.

The conventional process for treating zinc refinery residues is to recover Ga and Ge by leaching with sulfuric acid (Harbuck, 1993; Nishihama et al., 1999; Kul and Topkaya, 2008; Nusen et al., 2015; Zhang et al., 2016a). Unfortunately, this process has a major drawback as
the large amount of silica in the residue was leached to form a silica-germanium gel, which increases the co-precipitation of Ge. The most likely reason for this is the formation of mixed Ge-Si polymers which hydrolyze more readily than the Ge-polymers to form a germanium rich silica gel (Dutrizac et al., 1996; Efremov et al., 2002; Liang et al., 2009; Liu et al., 2016a). With longer reaction times, the losses of Ga and Ge increase and consequently, several efforts have been made to eliminate the negative influence of silicon on the recovery of Ga and Ge in sulfuric acid leaching. This has been achieved through the use of leaching assistant agents and surfactants (Liu et al., 2016a), pressure sulfuric acid leaching (Liu et al., 2016b), stepwise leaching (using sulfuric acid and hydrofluoric acid leaching sequentially) (Harbuck et al., 1989; Harbuck, 1993) and alkaline leaching (Torma, 1991; Lee et al., 1994). Nevertheless, some challenges still remain to the use of these processes on the industrial scale, for example, the enormous consumption of the leaching assistant reagents, high capital and operational costs of pressure acid leaching, the highly corrosive nature of hydrofluoric acid and subsequent difficult purification of the alkaline leaching due to the leaching of Si, Pb and Al. As a result, processing methods for high silica materials rich in Ga and Ge still need further development.

In this study, the oxalic acid leaching of high silica residue was carried out for Ga and Ge. Oxalic acid was utilized due to the fact that Ge is known to form stable complexes with di- and tricarboxylic acids, whereas Si forms much weaker complexes with these same ligands (Pokrovskii et al., 1998). In addition, oxalic is also widely used as metal precipitant (Chenakin et al., 2016; Kang et al., 2010; Ni et al., 2011) and thus, a selective leaching of the residues can be achieved. Sun et al. (2012), for example, have used oxalic acid leaching to treat spent lithium-ion batteries. Oxalate is introduced as leaching reagent meanwhile as
precipitant which leaches and precipitates cobalt from LiCoO$_2$ and CoO directly, thus the efficient separation of Li and Co can be achieved. Due to the strong complexation of ferric iron and oxalate, oxalic acid is also widely used in the removal of iron from quartz ores and these results have shown that the iron removal was above 90% using oxalic acid as leaching agent (Tuncuk et al., 2016; Lee et al., 2007; Ambikadevi et al., 2000). Szymczycha-Madeja (2011) and Liu et al. (2016c) investigated the separation and recovery of Mo and Ni in H$_2$C$_2$O$_4$ solution with H$_2$O$_2$. It was found that H$_2$O$_2$ can promote the leaching of Mo, and under optimal conditions, approximately 99.7% of Mo was leached, whilst only 1.6% of Ni was leached. Although there have been a number of previous studies into the effects of oxalic acid on the leaching of metals, such as Mo, Li, Ni and Fe, so far there is no report related to the recovery of gallium and germanium by an oxalic acid leaching process from zinc refinery residues.

The main purpose of the present work was to investigate the leaching behaviors of Ga and Ge in a solution containing both oxalic acid (as a chelating agent that is able to form soluble metal complexes) and hydrogen peroxide (as an oxidizing agent). Hydrogen peroxide was chosen not only because it is a strong oxidant but also as it is environmentally safe, with only water produced as a by-product during the sulfide oxidation process (Liu et al., 2016c; Zhang et al., 2016b; He et al., 2014; Szymczycha-Madeja, 2011).

Compared with the recovery of gallium and germanium in the sulfuric acid solution, the recovery of gallium and germanium in oxalic acid solution is more straightforward as both gallium and germanium can form stable complexes which can be effectively extracted by by tri(octyl-decyl) amine (N235). Furthermore, N235 is an amine extractant that is widely used for the extraction of complexes due to its low cost, high availability and high extraction efficiency (Fang et al., 2014; Yang et al., 2015; Chen
et al., 2016). Nevertheless, for the recovery of Ga and Ge, the removal of iron from the oxalate solution poses an additional challenge as Fe(III) is known to exist as a stable complex in oxalate solutions. The solvent extraction process (Demopoulos et al., 1993; Principe and Demopoulos, 2004, 2005; Sun et al., 2016; Mishra et al., 2010) and neutralization precipitation process (Liang et al., 2009) have previously been used for the removal of Fe(III) from different solutions, however, the drawbacks of these procedures can be the simultaneous co-extraction / precipitation / occlusion of Ga and Ge into the iron rich phases. As a result of this previous research, an ultrasound-assisted iron powder replacement method was proposed to remove the Fe in the form of FeC₂O₄·2H₂O from the oxalate solution.

In the purified oxalate solution, the extraction of gallium and germanium was conducted using tri(octyl-decyl) amine (N235) as an extractant and tri-butyl-phosphate (TBP) as a modifying agent. The separation of Ga and Ge from the organic phase was conducted using sulfuric acid and sodium hydroxide as stripping agents, respectively. In the current study the applicability of the proposed experimental flow sheet for the recovery of Ga and Ge and the behavior of the other elements of interest in each unit process were investigated in order verify the process validity.

2 Experimental

2.1 Materials

The zinc refinery residue investigated was obtained from the Danxia smelter in southern China and was used after drying with a vacuum drying oven (DZF-6050, Shanghai Yuezhong Instrument Co. Ltd, China), without any further pre-treatment. The contents of Ga, Ge and other main
elements of the interest present in the residue are shown in Table 1. The solid materials were firstly chemically dissolved, then the concentration of metal ions was analyzed by the methods listed in section 2.2.

As can be seen, the zinc refinery residue contains 0.266 wt.% Ga and 0.363 wt.% Ge respectively. In addition, whereas the total iron and Fe(II) content are 7.9 wt.% and 3.7 wt.% respectively. The silica content is observed to be relatively high (9.14 wt.%), which indicates that the leaching of gallium and germanium can be challenging.

The particle size of the zinc refinery residue was determined by a wet screen analysis, which showed that approximately 76 wt.% of the particles in the residue were < 45 μm (Table 2).

From analysis of the residue XRD pattern shown in Fig. 1, it can be clearly seen that there are several diffraction peaks of metallic zinc, zinc sulfate heptahydrate, zinc ferrite and zinc silicate, whereas no diffraction peaks of other components were observed because of their low contents or amorphous states. The small amounts of zinc ferrite (1.34 wt.%) and zinc silicate (2.87 wt.%) that are seen to exist in the residue probably originate from incomplete L/S (liquid-to-solid) separation after the pre-neutralization step and prior to the zinc powder replacement.

2.2 Characterization and analyses

The residue samples were measured by a Rigaku-TTRIII X-ray diffractometer (CuKα1, λ= 0.15406 nm) and the potential phases present in the sample have been identified by the software of MDI Jade 6.5. The elemental content in selected spots were measured by Electron Probe Microanalyzer (EPMA) with five wavelength dispersive X-ray spectrometers (WDS), a newly developed energy dispersive X-ray spectrometer (EDS) and the analysis was performed by a JEOL JXA-8230 instrument operated at 15 kV with a 10 nA beam current. The phase
compositions of the Ga and Ge in the zinc refinery residue were examined by use of a chemical selective dissolution method (Liu, 1992), and the analysis methods is mainly based on the dissolution behavior of gallium and germanium species in different solvents. The zinc concentrations in the leach residues and solutions were analyzed using ethylene diamine tetraacetic acid (EDTA) titration method. The SiO$_2$ concentrations in leach residues and solutions were determined using the silicon molybdenum blue spectrophotometry method, whereas the germanium concentrations in leach residues or solutions were examined by extraction separation - benzfluorenone spectrophotometry. The concentrations of Ga, Fe, Cu, As and Pb in leach residues or solutions were determined by a Thermo Electron IRIS Intrepid II XSP ICP-AES spectrometer. The Fe(Ⅱ) concentrations in the zinc refinery residue were analyzed by chemical dissolution - potassium dichromate method. In contrast, the Fe(Ⅲ) concentrations in leachate were analyzed by a precipitation separation - EDTA titration method (Yu et al., 2004). Additionally, the oxalate ions concentrations in the leachate were studied using the potassium permanganate titration method (Gong et al., 2006). Furthermore, the leach slurry was poured into a Φ100 Buchner funnel and filtered by vacuum suction using a SHZ-D(Ⅲ) water-ring pump with a vacuum degree of 0.07 MPa in order to observe the filtration performance according to the filtration rate.

2.3 Experimental procedure

The oxalic acid leaching experiments were conducted in 500 mL round-bottom flasks, with magnetic stirring at 300 rpm, respectively. Temperature was set and controlled by a water bath (DF-101S, Gongyi Yuhua Instrument Co., China). Iron removal experiments were initially performed for 100 mL of the resulting leachate in 250 mL conical flasks,
and the ultrasound time was set and controlled by an ultrasonic cleaner at 30 °C (KQ-2550E, Jiangsu Kunshan Ultrasonic Cleaning Instrument Co., China, 40 kHz, 200 W), then the sonification solution was heated at desired temperature by a water bath with mechanical stirring at 300 rpm. The amount of iron powder required is mainly determined by the content of Fe(III) concentration present in leachate after leaching. After the iron removal step, Ga and Ge present in the solution were extracted and enriched using solvent extraction. The experiments were performed using separation funnels (125 mL) by mixing an organic phase with the purified solution. Then, the loaded organic phase was subsequently stripped by sequential treatment by H$_2$SO$_4$ and NaOH. After each experiment, both residue and filtrate were subsequently analyzed and the investigated parameters utilized are displayed in Table 3. In the experiment, the organic phase consisted of tri(octyl-decyl) amine N235, tributyl phosphate (TBP) and sulfonated kerosene (all provided by Shanghai Rare-earth Chemical Co., Ltd., China) as the solvent for extraction. The iron powder and other chemical reagents used were of analytical grade, and deionized water was used throughout the experiments.

3 Results and Discussion

3.1 The occurrence forms of gallium and germanium

The sample of zinc refinery residue from the Danxia smelter was investigated with EPMA to determine the occurrence states of the gallium and germanium and the results are shown in Fig. 2 and Table 4. The analysis confirms the presence of zinc sulfate with a small quantity of Ga and Ge (analysis point 1). Additional points also showed an increased contents of Ga and Ge for the mixtures with iron, silica gels with zinc, lead, copper, calcium sulfates and metallic zinc, respectively.
Fig. 3 shows the distributions of Ge, Ga, Fe and SiO$_2$ in the zinc refinery residue and the images indicate that Ge and Ga mainly coexist in the presence of Fe and SiO$_2$.

Using the data listed in Table 4, the contents of Ga and Ge in the residue were determined as a function of the analogous Fe and Si contents by utilization of the following regression analyses presented in equations (1) – (4):

\[
\text{pct Ge} = 0.1132 + 0.0534 \text{ pct Si} \quad R=0.842 \quad (1) \\
\text{pct Ge} = 0.3061 + 0.0245 \text{ pct Fe} \quad R=0.597 \quad (2) \\
\text{pct Ga} = 0.2114 + 0.0321 \text{ pct Si} \quad R=0.298 \quad (3) \\
\text{pct Ga} = 0.1156 + 0.0061 \text{ pct Fe} \quad R=0.078 \quad (4)
\]

From the analysis it can be seen that there is a reasonably strong correlation between Ge and Si (as shown in Fig. 4) as well as between Ge and Fe. In contrast, the same analysis demonstrates only a relatively weak correlation between Ga and Si and between Ga and Fe. From these results, it can be suggested that Ge is mainly co-precipitated with iron gel and silica gel in the form of MeO·GeO$_2$, while Ga is mainly displaced by zinc powder into the metallic form and then oxidized to Ga$_2$O$_3$.

The chemical selective dissolution method was utilized in order to more accurately analyze the phases of gallium and germanium present in the residue. The results outlined in Table 5 provide additional confirmation of the results from EPMA that showed that gallium and germanium primarily exist in the forms of dissociative or combined oxides.

### 3.2 Oxalic acid leaching

As can be seen from Table 5, in addition to the dissociative or combined oxides, germanium also occurred as both metallic germanium and germanium sulfide. The small amounts of metallic germanium and
germanium sulfide probably originate from unreacted zinc sulfide concentrate as this contains Ga and Ge due to incomplete L/S (liquid-to-solid) separation after the oxygen pressure leaching. Alternatively, the source can also be the conversion of sulfide in acidic solution because zinc calcine contains a small amount of sulfide. Based on the dissolution properties of germanium species, metallic germanium and sulfide are significantly more stable than other germanium species, which gives rise to the difficulties encountered when leaching germanium. In order to solve this problem, H₂O₂ was introduced as an oxidizing agent to convert the insoluble germanium species into more soluble forms. Furthermore, oxalic acid was utilized to improve both the leaching of Ga and Ge and the filtration performance of the leaching slurry.

3.2.1 Effect of oxalic acid concentration

The effects of oxalic acid concentration on the leaching efficiencies of Ga, Ge, Fe, Zn, Cu and Si are shown in Fig. 5. The results clearly indicate that the extraction of Ga, Ge, Fe and Zn are strongly dependent on the increase in the H₂C₂O₄ concentration, but that the extraction of Cu and Si only change slightly. The main reaction in the leaching are shown in Table 6. The extraction of Ga and Ge increased from 75.04 to 98.89% and 37.33 to 94.19% respectively as H₂C₂O₄ concentration increased from 40 g/L to 110 g/L. Conversely, further increases in H₂C₂O₄ concentration resulted in a slight decrease in the leaching of Ga and Ge probably as a result of increased Si leaching (from 0.38 to 0.56%), which forms a silica gel that absorbs some Ga and Ge (Liang et al., 2009; Liu et al., 2016b), simultaneously resulting in a slight decrease in the slurry filtration performance (860.62 to 790.86 L/m²·h). Under the same conditions, the leaching of iron increased slightly from 14.38 to 32.33% with an increase in the H₂C₂O₄ concentration from 40 g/L to 130 g/L. The
low leaching efficiency of iron over the investigated oxalic acid concentration range can be attributed to the simultaneous precipitation reaction of ferrous ions and oxalic acid when the complexation reaction of Fe(III) and H$_2$C$_2$O$_4$ occurs, which leads to the formation of FeC$_2$O$_4$·2H$_2$O. Ferrous ions in the solution can originate from the zinc refinery residues or from the reduction of Fe(III) (Taxiarchou et al., 1997; Veglio et al., 1999). In contrast, the leaching of both Zn and Cu decreased, when the acid concentration was increased from 40 g/L to 130 g/L - Zn decreased from 13.29 % to 0.26 %, whereas Cu changed from 2.31% to 0.65 %. This finding results from the formation of stable precipitates of ZnC$_2$O$_4$·2H$_2$O and CuC$_2$O$_4$ as demonstrated by the XRD results of the leaching residues shown in Fig. 6.

3.2.2 Effect of leaching time

The effect of leaching time on the recovery of Ga, Ge, Fe, Zn, Cu and Si are presented in Fig. 7. It can be seen from the results that the leaching of Ga and Ge were efficient and fast with 98.89 % of the Ga and 94.19 % of the Ge extracted within 30 min of the leaching with only a minor further increase observed when leaching time was further extended from 30 to 60 min. The leaching of Fe initially increased as the leaching time changed from 10 to 20 min (14.55 to 32.15%), however, this was followed by a reduction in leaching efficiency from 32.15 to 24.14% with increasing leaching time from 20 min to 40 min. Nevertheless, further increases in leaching time up to 60 mins resulted in further enhancement of the amounts of Fe leached. The reason for this observed variation in the Fe behavior is suggested to be attributed to the formation of FeC$_2$O$_4$·2H$_2$O between 20 to 40 min, and the oxidation of the Fe(II) in the residues by H$_2$O$_2$ and air after 40 min. In contrast to the leaching of Fe, in the leaching of Zn and Cu after approximately 20 min, a continuous decrease can be observed. The leaching of Si varies only slightly between
10 to 60 min, which results in an improved leach slurry filtration performance in oxalic acid when compared to that obtained with sulfuric acid (Liu et al., 2016a; Liu et al., 2016b).

3.2.3 Effect of the L/S ratio

Fig. 8 shows the effects of the liquid-to-solid (L/S) ratio on the leaching of Ga, Ge, Fe, Cu, Zn and Si. The results indicate that the leaching of Ga and Ge increased from 84.21 to 98.89% and 80.95 to 94.19%, respectively, as the L/S ratio increased from 4 to 8. This change occurs as a result the presence of gallium and germanium in the leach solution in their oxyacid forms, which can combined with the cations present (Zn$^{2+}$, Cu$^{2+}$ and Fe$^{2+}$) and precipitate as gallate and germanate. Hence, the leaching of Ga and Ge are controlled by the solubility of gallate and germanate. In addition, the increase in the liquid to solid ratio, means that consequently increase H$_2$C$_2$O$_4$ and H$_2$O$_2$ amounts in the leaching solution to participate in the reaction at a preset leaching time (Liu et al., 2016c), thus a higher L/S ratio gives rise to enhanced levels of gallium and germanium leaching.

The leaching of Fe increased slightly when the L/S increased from 4 to 12 (12.69 to 33.95%), whereas the leaching of Zn and Cu decreased slightly (4.35 to 0.27%, 2.98 to 0.35%). Under the same conditions, the leaching of Si was observed to increase from 0.12 to 1.2% (with L/S from 4 to 12), which can lead to the deterioration of the filtration performance of the leach slurry (860.62 to 698.45 L/m$^2$·h) due to the formation of silica gel at higher L/S ratios (Hua et al., 2002; Harbuck, 1993; Bodas, 1996). In addition, as the concentration of Ga and Ge in the leaching solution decreased as the L/S ratio increased, an optimum L/S ratio of 8 was chosen for the subsequent experiments.
3.2.4 Effect of temperature

The leaching results obtained at different temperatures (20 - 60 °C) are shown in Fig. 9. As can be observed, the leaching of both Ga and Ge increased with increasing leaching temperature with a maximum of 98.89% (Ga) and 94.19% (Ge) reached at 40 °C, respectively. Further increases in the leaching temperature resulted in a decrease in gallium and germanium leaching, which can be attributed to the thermolabile nature of H₂O₂ (Arvin and Pedersen, 2015; Zhang et al., 2016b), which causes the oxidation efficiency of H₂O₂ on insoluble substances, like Ge, GeS₂, GeS and Ga₃S₃, to decrease with increased temperature. Under the same conditions, the leaching of Fe increased slightly when the temperature increased from 20 °C to 60 °C (21.75 to 29.15%), whereas the leaching of Zn, Cu and Si were all below 1% over the same range. Based on these results in this work, it can be considered that the optimal temperature for leaching with oxalate and hydrogen peroxide is around 40 °C.

3.2.5 Effect of H₂O₂ concentration

The effect of H₂O₂ concentration was examined by varying the H₂O₂ concentration from 0 mol/L to 0.2 mol/L, as shown in Fig. 10. The results indicate that the extraction of Ge increased gradually with the increase of H₂O₂ concentration. The leaching of Ge reached yields of 98.86% when the H₂O₂ concentration reached 0.12 mol/L, which equates to an improvement of approximately 15% when compared to direct H₂C₂O₄ leaching method without H₂O₂. Nonetheless, further increases in the H₂O₂ concentration, only led to a slight further acceleration in the leaching of Ge. On the other hand, the leaching of Ga only changed slightly with an increase in the H₂O₂ concentration from 0 mol/L to 0.2 mol/L (from 95.78 to 99.86%). The main reason for the measured increases of Ga and Ge leaching are attributed to the oxidation reactions of H₂O₂ with insoluble materials (Ge, GeS₂, GeS and Ga₃S₃) and the complexation reactions of
H_2C_2O_4 with Ga and Ge. The main oxidation reactions are expressed in Eqs. (13) - (17), and the changes in the corresponding Gibbs free energy values are listed in Table 7.

\[
\begin{align*}
\text{GeS}_2 + 8\text{H}_2\text{O}_2 &= \text{GeO}_2 + 2\text{SO}_4^{2-} + 4\text{H}^+ + 6\text{H}_2\text{O} \quad (13) \\
\text{Ge} + 2\text{H}_2\text{O}_2 &= \text{GeO}_2 + 2\text{H}_2\text{O} \quad (14) \\
\text{GeS} + 4\text{H}_2\text{O}_2 &= \text{GeO} + \text{SO}_4^{2-} + 2\text{H}^+ + 3\text{H}_2\text{O} \quad (15) \\
\text{Ga}_2\text{S}_3 + 12\text{H}_2\text{O}_2 &= \text{Ga}_2\text{O}_3 + 3\text{SO}_4^{2-} + 6\text{H}^+ + 9\text{H}_2\text{O} \quad (16) \\
2\text{Ga} + 3\text{H}_2\text{O}_2 &= \text{Ga}_2\text{O}_3 + 3\text{H}_2\text{O} \quad (17)
\end{align*}
\]

As shown in Table 7, the change in the Gibbs free energy for each reaction is negative at 40 °C, which suggests that the insoluble elements and sulfides of gallium and germanium are thermodynamically likely to transform into the corresponding soluble oxides when H_2O_2 is introduced into the system, which effectively promotes gallium and germanium leaching. Due to the fact that this process involves the use of both a complexing agent and oxidant, part of Fe(III) in the leaching residues will also be leached through the formation of iron complexes. This is evident from the results outlined in Fig. 10, which show that the leaching of Fe increases from 18.15% to 31.08% as the H_2O_2 increased from 0 mol/L to 0.2 mol/L. Under the same conditions, the leaching of Cu, Zn and Si all change only slightly with an increase in the H_2O_2 concentration from 0 mol/L to 0.2 mol/L, with changes from 0.73 to 0.81% (Cu), 0.14 to 0.41% (Zn) and 0.39 to 0.42% (Si), respectively.

### 3.3 Removal of iron from oxalate solution

Based on the results outlined above, the optimal leaching conditions were determined to be a H_2C_2O_4 concentration of 110 g/L and a H_2O_2 concentration of 0.12 mol/L with a L/S ratio of 8 at 40 °C for 30 min.
Under these optimum conditions, the leaching efficiencies of Ga, Ge, Fe, Zn, Cu and Si attained were 99.32%, 98.86%, 30.25%, 0.30%, 0.82% and 0.43%, respectively, with a composition of 0.330 g/L Ga, 0.447 g/L Ge, 3.010 g/L Fe, 0.092 g/L Zn, 0.056 g/L Cu and 0.023 g/L Si. The XRD of leaching residues following this treatment showed that the main phases present in the leach residues were ZnC_2O_4·2H_2O, FeC_2O_4·2H_2O, CuC_2O_4 and SiO_2, and the contents of main elements in the leaching residues are shown in Table 8. As was observed, more than 30% of the Fe was leached in the forms of iron complexes during the oxalic acid leaching, and this can have a detrimental impact on both the extraction of Ga and Ge and the recycling efficiency of the extractant in the subsequent Ga and Ge solvent extraction processes. In order to address this challenge, the ultrasound-assisted iron powder replacement process was adopted.

The effects of main factors (Fe/Fe(III) molar ratio, ultrasound time, water bath reaction time and water bath temperature) on the iron removal were examined with an ultrasound power of 150 W. The results shown in Table 9 indicated that the iron removal ratio is strongly depended on the Fe/Fe(III) molar ratio and ultrasound time. The iron removal ratio increased from 78.69 to 98.31% as the Fe/Fe(III) ratio increased from 2 to 6, whereas, the iron removal ratio changed only slightly with a further increase from 6 to 8 (98.31 to 99.01%). More critically however, the loss of Ga and Ge increased from 1.08 to 2.55% and 0.68 to 1.58%, respectively, when the Fe/Fe(III) ratio was increased from 6 to 8. As a result, the optimum Fe/Fe(III) ratio selected to be 6. The main reaction is outlined in Eq. (18):

$$2\text{Fe(C}_2\text{O}_4\text{)}^+ + \text{Fe} + \text{C}_2\text{O}_4^{2-} + 2\text{H}_2\text{O} = 3\text{FeC}_2\text{O}_4 \cdot 2\text{H}_2\text{O} \quad \log K_{(298.15)} = 19.86 \quad (18)$$

In addition, Table 9 also shows that the sonification is necessary for the highly efficiency iron removal. For example, the iron removal ratio can
increase to 98.31% when an ultrasound power of 150 W is applied for 10 min, which equates to more than approximately 50% improvement in the iron removal when compared to that achieved without ultrasonic treatment. An increase in the water bath reaction time and temperature also aid iron removal and the crystallization of the FeC$_2$O$_4$·2H$_2$O.

The optimum conditions for the maximum removal of 98.31% of the Fe and minimum losses of 1.08% of the Ge and 0.68% of the Ga were determined to be as follows: Fe(iron powder)/Fe(III) molar ratio of 6, ultrasound power of 150 W; ultrasound time of 10 min; ultrasound temperature of 30 °C; water bath temperature of 50 °C and water bath time of 1 h following the application of ultrasound. The main phases present in the residue were FeC$_2$O$_4$·2H$_2$O (Fig. 11), and a purified solution containing 0.327 g/L Ga, 0.442 g/L Ge, 0.050 g/L Fe, 0.089 g/L Zn, 0.014 g/L Cu and 0.015 g/L Si.

3.4 Separation of Ga and Ge from purified solution

In order to recovery and enrich the Ga and Ge effectively, a solvent extraction method was carried out using a mixed extractant that comprised of 20% (v/v) N235, 10% (v/v) TBP and 70% (v/v) sulfonated kerosene. The optimum conditions determined for the solvent extraction were: a contact time of 15 min, temperature of 25 °C and an O:A ratio of 1:4. Under these optimum conditions, 99.36 % of the Ga and 99.89% of the Ge could be extracted and a loaded organic phase that contained 1.308 g/L Ga and 1.766 g/L Ge was obtained. The separation of Ga and Ge from the organic phase was conducted using sequential treatment steps with H$_2$SO$_4$ and NaOH solution.

Firstly, 98.91% of the Ga and 0.02% of the Ge in the loaded organic phase were stripped respectively by H$_2$SO$_4$ (2 mol/L) for 15 min at 30 °C with an A:O ratio of 1:1, and a stripping liquor containing 1.290 g/L Ga
and 0.004 g/L Ge was obtained. After Ga stripping, 99.21% of the Ge and 0.52% of the Ga in the loaded organic phase were then stripped using NaOH (4 mol/L) for 15 min at 40 °C with an A:O ratio of 1:1 resulting in a stripping liquor that contained 0.006 g/L Ga and 1.751 g/L Ge, respectively.

According to the results obtained in the experiments presented above, the proposed flow sheet for treatment of zinc refinery residues containing high silica and iron is shown in Fig.12.

4 Conclusions

On the basis of the comprehensive characterization of materials, in this paper a hydrometallurgical process for treating a Ga and Ge rich zinc refinery residue was investigated. It was found that Ga and Ge in the zinc refinery residue mainly exist as dissociative or combined oxides that correlate with the presence of iron and silica. Under leaching conditions that comprised of an H_2C_2O_4 concentration of 110 g/L, H_2O_2 concentration of 0.12 mol/L, a leaching time of 30 min, an L/S ratio of 8 and a temperature of 40 °C, 99.32% of the Ga, 98.86% of the Ge and 30.25% of the Fe were leached out. In contrast, the leaching of Zn, Cu and Si was only 0.30%, 0.82% and 0.43%, respectively. Purification by the ultrasound-assisted iron powder replacement process allowed 98.31% of the iron to be removed in the form of FeC_2O_4·2H_2O, whilst the concurrent losses of Ga and Ge were found to be 1.08% and 0.68%, respectively. After purification, 99.36 % of the Ga and 99.89% of the Ge were extracted by N235 : TBP : kerosene. Finally, 98.91% of the Ga and 99.21% of the Ge loaded in the organic phase were stripped into different solutions using sequential treatments of 2 mol/L H_2SO_4 and 4 mol/L NaOH. The resultant pregnant solutions obtained had concentrations of 1.290 g/L Ga (H_2SO_4) and 1.751 g/L Ge (NaOH). Pure gallium can be
obtained by electrolysis after adjusting the alkalinity of the enriched gallium solution with NaOH, whereas the germanium can be further recovered by chlorinated distillation process from the products that originate from neutralization precipitation of enriched germanium solution. In addition, the effective recovery of valuable metals (Zn, Fe, Cu, Pb) from leaching residues can be realized by the combination of rotary kiln calcination and hydrometallurgical treatment processes. The results detailed in paper demonstrate that almost complete separation and recovery of gallium and germanium from the zinc refinery residue could be obtained using this technology.

Acknowledgements

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References


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iron(III) from sulphuric acid solutions with mixed kelex 100-alkyl phosphorus acid extractants, Solvent Extr. Ion Exch. 11 (1), 67–89.


Sun, L., Qiu, K.Q., 2012. Organic oxalate as leachant and precipitant for the recovery of valuable metals from spent lithium-ion batteries. Waste Management 32(8), 1575–1582.


silica sands production by oxalic acid leaching. Industrial and Engineering Chemistry Research 38, 4443–4448.
Fig. 1. X-ray diffraction pattern of the zinc refinery residue.
Fig. 2. Backscattered electron image of the zinc refinery residue.
Fig. 3. Distribution maps of the Ge, Ga, Fe and Si in the zinc refinery residue.
Fig. 4. The correlation of Si content and the amount of Ge present in the zinc refinery residue based on EPMA analysis in Table 4.
Fig. 5. Effect of oxalic acid concentration on the recovery of Ga, Ge, Fe, Cu, Zn and Si to the leach solution during zinc refinery residue leaching (T = 40 °C; L/S = 8; t = 30 min; \([H_2O_2] = 0.08\) mol/L).
Fig. 6. X-ray diffraction pattern of the leaching residue after oxalic acid leaching of zinc refinery residue ([H$_2$C$_2$O$_4$] = 110 g/L; T = 40 °C; L/S = 8; t = 30 min; [H$_2$O$_2$] = 0.08 mol/L).
Fig. 7. Effect of leaching time on the recovery of Ga, Ge, Fe, Cu, Zn and Si to the leach solution during zinc refinery residue leaching ($[H_2C_2O_4] = 110 \text{ g/L}; \text{L/S} = 8; \text{T} = 40 \degree \text{C}; [H_2O_2] = 0.08 \text{ mol/L}$).
Fig. 8. Effect of L/S ratio on the recovery of Ga, Ge, Fe, Cu, Zn and Si to the leach solution during zinc refinery residue leaching ($[H_2C_2O_4] = 110$ g/L; $T = 40$ °C; $t = 30$ min; $[H_2O_2] = 0.08$ mol/L).
Fig. 9. Effect of temperature on the recovery of Ga, Ge, Fe, Cu, Zn and Si to the leach solution during zinc refinery residue leaching ($[\text{H}_2\text{C}_2\text{O}_4] = 110$ g/L; L/S = 8; $t = 30$ min; $[\text{H}_2\text{O}_2] = 0.08$ mol/L).
Fig. 10. Effect of $\text{H}_2\text{O}_2$ concentration on the recovery of Ga, Ge, Fe, Cu, Zn and Si to the leach solution during zinc refinery residue leaching ($[\text{H}_2\text{C}_2\text{O}_4] = 110 \text{ g/L}; \text{L/S} = 8; \text{T} = 40{^\circ}\text{C}; \text{t} = 30 \text{ min}$).
Fig. 11. X-ray diffraction pattern of the residue after iron removal (Fe/Fe(III) = 6; Ultrasound power = 150 W; Ultrasound time = 10 min; T = 50 °C; t = 1 h).

* FeC$_2$O$_4$•2H$_2$O
**Fig. 12.** Process flow sheet for the recovery Ga and Ge from zinc refinery residues.
Table 1. Contents of the major component of interest in the zinc refinery residue. (wt.%).

<table>
<thead>
<tr>
<th></th>
<th>Zn</th>
<th>Cu</th>
<th>Fe</th>
<th>Fe(II)</th>
<th>SiO₂</th>
<th>Ga</th>
<th>Ge</th>
<th>Pb</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>24.45</td>
<td>5.55</td>
<td>7.88</td>
<td>3.74</td>
<td>9.14</td>
<td>0.266</td>
<td>0.362</td>
<td>0.46</td>
</tr>
</tbody>
</table>
Table 2. Particle size distribution of the zinc refinery residue by wet screen analysis.

<table>
<thead>
<tr>
<th>Particle size/μm</th>
<th>&lt;45</th>
<th>45-60</th>
<th>60-93</th>
<th>93-150</th>
<th>&gt;150</th>
</tr>
</thead>
<tbody>
<tr>
<td>wt.%</td>
<td>75.81</td>
<td>4.81</td>
<td>4.66</td>
<td>10.83</td>
<td>3.89</td>
</tr>
</tbody>
</table>
Table 3. Main parameters investigated in the experiments.

<table>
<thead>
<tr>
<th>Experiment</th>
<th>Investigated parameters</th>
</tr>
</thead>
<tbody>
<tr>
<td>Oxalic acid leaching stage</td>
<td>L/S (mL/g): 4-12; Temperature (°C): 30-60; H$_2$C$_2$O$_4$ (g/L): 40-130; H$_2$O$_2$ (mol/L): 0-0.2; Reaction time (min): 10-60</td>
</tr>
<tr>
<td>Iron removal</td>
<td>Fe/Fe(III) mole ratio: 2-8; Ultrasound (200 W) time (min): 0-15;</td>
</tr>
<tr>
<td></td>
<td>Water bath temperature (°C): 30-70; Water bath time (h): 0.5-2</td>
</tr>
<tr>
<td>Solvent extraction</td>
<td>Reaction time (min): 5-20; Temperature (°C): 15-35; N$_2$35(v/v): 5-10%; Organic/aqueous phase ratio(v/v)(O:A): 2:1-1:6</td>
</tr>
<tr>
<td>Stripping</td>
<td>Reaction time (min): 10-30; Temperature (°C): 25-50; H$_2$SO$_4$ (mol/L): 1-4; NaOH (mol/L): 2-6; A:O: 1:4-2:1</td>
</tr>
</tbody>
</table>
Table 4. EPMA results of the analysis points of the zinc refinery residues displayed in Figure 2 (wt.%).

<table>
<thead>
<tr>
<th>Spot</th>
<th>O</th>
<th>Ge</th>
<th>S</th>
<th>Ga</th>
<th>Zn</th>
<th>Fe</th>
<th>Cu</th>
<th>Si</th>
<th>Pb</th>
<th>Ca</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>39.703</td>
<td>0.073</td>
<td>10.587</td>
<td>0.041</td>
<td>22.308</td>
<td>1.702</td>
<td>3.119</td>
<td>2.682</td>
<td>0.256</td>
<td>0.060</td>
</tr>
<tr>
<td>2</td>
<td>39.922</td>
<td>0.838</td>
<td>2.708</td>
<td>0.431</td>
<td>21.016</td>
<td>5.868</td>
<td>5.564</td>
<td>6.476</td>
<td>1.28</td>
<td>ND</td>
</tr>
<tr>
<td>3</td>
<td>41.632</td>
<td>0.813</td>
<td>2.590</td>
<td>0.312</td>
<td>12.867</td>
<td>5.947</td>
<td>2.728</td>
<td>12.384</td>
<td>0.283</td>
<td>5.450</td>
</tr>
<tr>
<td>4</td>
<td>41.476</td>
<td>0.407</td>
<td>4.259</td>
<td>0.237</td>
<td>24.283</td>
<td>5.475</td>
<td>5.523</td>
<td>2.253</td>
<td>0.319</td>
<td>0.467</td>
</tr>
<tr>
<td>5</td>
<td>55.199</td>
<td>0.506</td>
<td>3.204</td>
<td>0.482</td>
<td>14.014</td>
<td>4.526</td>
<td>5.685</td>
<td>6.164</td>
<td>0.723</td>
<td>10.45</td>
</tr>
<tr>
<td>6</td>
<td>27.399</td>
<td>0.104</td>
<td>0.492</td>
<td>0.159</td>
<td>16.792</td>
<td>41.075</td>
<td>1.229</td>
<td>0.483</td>
<td>0.069</td>
<td>ND</td>
</tr>
<tr>
<td>7</td>
<td>46.982</td>
<td>0.046</td>
<td>10.934</td>
<td>0.500</td>
<td>22.665</td>
<td>2.547</td>
<td>3.701</td>
<td>1.245</td>
<td>0.472</td>
<td>14.050</td>
</tr>
<tr>
<td>8</td>
<td>40.975</td>
<td>0.725</td>
<td>1.348</td>
<td>0.286</td>
<td>13.545</td>
<td>5.245</td>
<td>1.756</td>
<td>14.548</td>
<td>0.210</td>
<td>3.145</td>
</tr>
<tr>
<td>9</td>
<td>46.145</td>
<td>0.564</td>
<td>6.145</td>
<td>0.524</td>
<td>24.145</td>
<td>2.156</td>
<td>5.482</td>
<td>13.245</td>
<td>0.275</td>
<td>0.896</td>
</tr>
<tr>
<td>10</td>
<td>47.176</td>
<td>0.542</td>
<td>4.235</td>
<td>0.422</td>
<td>23.172</td>
<td>3.458</td>
<td>4.467</td>
<td>9.789</td>
<td>0.315</td>
<td>3.456</td>
</tr>
<tr>
<td>11</td>
<td>28.455</td>
<td>0.115</td>
<td>0.385</td>
<td>0.162</td>
<td>17.042</td>
<td>42.025</td>
<td>0.768</td>
<td>0.445</td>
<td>0.104</td>
<td>0.135</td>
</tr>
<tr>
<td>12</td>
<td>50.324</td>
<td>0.465</td>
<td>4.865</td>
<td>0.452</td>
<td>16.415</td>
<td>2.576</td>
<td>5.425</td>
<td>6.784</td>
<td>0.113</td>
<td>2.745</td>
</tr>
<tr>
<td>13</td>
<td>39.017</td>
<td>0.135</td>
<td>0.296</td>
<td>0.344</td>
<td>20.319</td>
<td>37.530</td>
<td>0.703</td>
<td>0.549</td>
<td>0.066</td>
<td>1.087</td>
</tr>
<tr>
<td>14</td>
<td>42.325</td>
<td>0.856</td>
<td>3.096</td>
<td>0.295</td>
<td>12.067</td>
<td>5.846</td>
<td>2.708</td>
<td>15.384</td>
<td>0.243</td>
<td>2.645</td>
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<tr>
<td>15</td>
<td>46.785</td>
<td>0.607</td>
<td>4.651</td>
<td>0.197</td>
<td>25.086</td>
<td>5.145</td>
<td>5.420</td>
<td>11.253</td>
<td>0.319</td>
<td>0.467</td>
</tr>
</tbody>
</table>

ND – not detected at a minimum detection limit of 300 ppm.
Table 5. Phase composition of gallium and germanium in the zinc refinery residue.

<table>
<thead>
<tr>
<th>Germanium phases</th>
<th>wt.%</th>
<th>Distribution/%</th>
<th>Gallium phases</th>
<th>wt.%</th>
<th>Distribution/%</th>
</tr>
</thead>
<tbody>
<tr>
<td>MeO·GeO&lt;sub&gt;2&lt;/sub&gt;</td>
<td>0.222</td>
<td>61.32</td>
<td>MeO·Ga&lt;sub&gt;2&lt;/sub&gt;O&lt;sub&gt;3&lt;/sub&gt;</td>
<td>0.076</td>
<td>28.58</td>
</tr>
<tr>
<td>GeO+GeO&lt;sub&gt;2&lt;/sub&gt;</td>
<td>0.081</td>
<td>22.38</td>
<td>Ga&lt;sub&gt;2&lt;/sub&gt;O&lt;sub&gt;3&lt;/sub&gt;</td>
<td>0.187</td>
<td>70.30</td>
</tr>
<tr>
<td>Ge+GeS&lt;sub&gt;2&lt;/sub&gt;+GeS</td>
<td>0.059</td>
<td>16.30</td>
<td>Ga+Ga&lt;sub&gt;2&lt;/sub&gt;S&lt;sub&gt;3&lt;/sub&gt;</td>
<td>0.003</td>
<td>1.12</td>
</tr>
<tr>
<td>Total</td>
<td>0.362</td>
<td>100</td>
<td>Total</td>
<td>0.266</td>
<td>100</td>
</tr>
</tbody>
</table>
### Table 6. The main reaction and corresponding equilibrium constants in oxalate leaching stage (298.15 K).

<table>
<thead>
<tr>
<th>No.</th>
<th>Reaction</th>
<th>logK</th>
<th>References</th>
</tr>
</thead>
<tbody>
<tr>
<td>5</td>
<td>$\text{Ga}_2\text{O}_3 + 6\text{HC}_2\text{O}_4^- = 2\text{Ga(C}_2\text{O}_4)_2^{2-} + 3\text{H}_2\text{O}$</td>
<td>9.12</td>
<td>Weast (1976)</td>
</tr>
<tr>
<td></td>
<td></td>
<td></td>
<td>HSC 7.0</td>
</tr>
<tr>
<td>6</td>
<td>$\text{GeO}_2 + 2\text{HC}_2\text{O}_4^- = \text{Ge(OH)}_2\text{(C}_2\text{O}_4)_2^{2-}$</td>
<td>2.95</td>
<td>Everest (1955)</td>
</tr>
<tr>
<td></td>
<td></td>
<td></td>
<td>Pokrovskii (1998)</td>
</tr>
<tr>
<td>7</td>
<td>$\text{SiO}_2 + \text{HC}_2\text{O}_4^- + \text{H}^+ = \text{Si(OH)}_2\text{C}_2\text{O}_4$</td>
<td>-16.64</td>
<td>Bennett (1991)</td>
</tr>
<tr>
<td></td>
<td></td>
<td></td>
<td>Pokrovskii (1998)</td>
</tr>
<tr>
<td>8</td>
<td>$\text{Fe}^{2+} + \text{C}_2\text{O}_4^{2-} + 2\text{H}_2\text{O} = \text{FeC}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$</td>
<td>6.47</td>
<td>Weast (1976)</td>
</tr>
<tr>
<td>9</td>
<td>$\text{Fe}_2\text{O}_3 + 2\text{HC}_2\text{O}_4^- + 4\text{H}^+ = 2\text{Fe(C}_2\text{O}_4)_2^{+} + 3\text{H}_2\text{O}$</td>
<td>10.50</td>
<td>Weast (1976)</td>
</tr>
<tr>
<td>10</td>
<td>$\text{ZnFe}_2\text{O}_4 + 3\text{HC}_2\text{O}_4^- + 5\text{H}^+ = \text{ZnC}_2\text{O}_4 \cdot 2\text{H}_2\text{O} + 2\text{Fe(C}_2\text{O}_4)_2^{+} + 3\text{H}_2\text{O}$</td>
<td>14.30</td>
<td>Weast (1976)</td>
</tr>
<tr>
<td>11</td>
<td>$\text{ZnO} + \text{HC}_2\text{O}_4^- + \text{H}^+ + \text{H}_2\text{O} = \text{ZnC}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$</td>
<td>15.70</td>
<td>Weast (1976)</td>
</tr>
<tr>
<td>12</td>
<td>$\text{CuO} + \text{HC}_2\text{O}_4^- + \text{H}^+ = \text{CuC}_2\text{O}_4 + \text{H}_2\text{O}$</td>
<td>8.89</td>
<td>Weast (1976)</td>
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</table>
Table 7. Gibbs free energy change of oxidation reaction involved with H$_2$O$_2$.

<table>
<thead>
<tr>
<th>Equation No.</th>
<th>The values of $\Delta G^\circ$ for the reactions (313.15 K), (kJ/mol)</th>
</tr>
</thead>
<tbody>
<tr>
<td>(13)</td>
<td>-2228.870</td>
</tr>
<tr>
<td>(14)</td>
<td>-725.628</td>
</tr>
<tr>
<td>(15)</td>
<td>-1037.950</td>
</tr>
<tr>
<td>(16)</td>
<td>-3237.767</td>
</tr>
<tr>
<td>(17)</td>
<td>-653.126</td>
</tr>
</tbody>
</table>

The values of $\Delta G^\circ$ were calculated using HSC Chemistry® 7.0 software.
Table 8. Contents of the major component of interest in the leaching residue (wt.%).

<table>
<thead>
<tr>
<th></th>
<th>Zn</th>
<th>Cu</th>
<th>Fe</th>
<th>SiO₂</th>
<th>Ga</th>
<th>Ge</th>
<th>Pb</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>19.37</td>
<td>4.62</td>
<td>4.14</td>
<td>16.14</td>
<td>0.005</td>
<td>0.012</td>
<td>0.30</td>
</tr>
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</table>

<table>
<thead>
<tr>
<th></th>
<th>Ca</th>
<th>Co</th>
<th>Ni</th>
<th>Al</th>
<th>As</th>
<th>Sn</th>
<th>S</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>0.62</td>
<td>0.16</td>
<td>0.04</td>
<td>0.48</td>
<td>0.11</td>
<td>0.30</td>
<td>1.46</td>
</tr>
</tbody>
</table>

Table 9. Effect of ultrasound time, Fe/Fe(III) molar ratio, reaction temperature and reaction time on iron removal.

<table>
<thead>
<tr>
<th>Factors</th>
<th>Fe (%)</th>
<th>Ge (%)</th>
<th>Ga (%)</th>
</tr>
</thead>
<tbody>
<tr>
<td>Fe/Fe(III) molar ratio</td>
<td>2</td>
<td>78.69</td>
<td>0.59</td>
</tr>
<tr>
<td>(Ultrasound time:10 min; 50°C; 1h)</td>
<td>4</td>
<td>90.30</td>
<td>0.91</td>
</tr>
<tr>
<td></td>
<td>6</td>
<td>98.31</td>
<td>1.08</td>
</tr>
<tr>
<td></td>
<td>8</td>
<td>99.01</td>
<td>2.55</td>
</tr>
<tr>
<td>Ultrasound time (min)</td>
<td>0 min</td>
<td>50.32</td>
<td>0.55</td>
</tr>
<tr>
<td>(Fe/Fe(III): 6; 50°C; 1h)</td>
<td>5 min</td>
<td>88.30</td>
<td>0.61</td>
</tr>
<tr>
<td></td>
<td>10 min</td>
<td>98.31</td>
<td>1.08</td>
</tr>
<tr>
<td></td>
<td>15 min</td>
<td>98.56</td>
<td>1.10</td>
</tr>
<tr>
<td>Water bath temperature (°C)</td>
<td>30 °C</td>
<td>90.85</td>
<td>0.56</td>
</tr>
<tr>
<td>(Ultrasound time:10 min; Fe/Fe(III): 6; 1h)</td>
<td>50 °C</td>
<td>98.31</td>
<td>1.08</td>
</tr>
<tr>
<td></td>
<td>70 °C</td>
<td>98.71</td>
<td>2.17</td>
</tr>
<tr>
<td>Water bath time (h)</td>
<td>0.5 h</td>
<td>92.75</td>
<td>0.68</td>
</tr>
<tr>
<td>(Ultrasound time:10 min; Fe/Fe(III): 6; 50°C)</td>
<td>1 h</td>
<td>98.31</td>
<td>1.08</td>
</tr>
<tr>
<td></td>
<td>1.5 h</td>
<td>98.90</td>
<td>1.24</td>
</tr>
</tbody>
</table>
Research highlights

(1) Ga and Ge exist as dissociative or combined oxides correlated with Fe and Si.
(2) Ga and Ge were selectively leached in H₂C₂O₄ solution with H₂O₂.
(3) Ultrasound-assisted replacement can effectively remove iron.
(4) Ga and Ge were almost totally extracted by N235:TBP:kerosene.